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M HCl  $[H^+] = 0.01M$  pH =

$-\log(0.01) = 2$  2. 0.0010 M NaOH

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$[\text{OH}^-] = 0.0010\text{M}$   $\text{pOH} =$   
 $-\log(0.010) = 3$       $\text{pH} = 11.3$ .  
0.050 M NaOH

pH and pOH CONTINUED -  
Newbury Park High School  
pH/pOH. pH and pOH are a  
measure of how acidic or how

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basic a solution is and for most solutions, is measured on a scale of 0 to 14. The lower the number, the more acidic the solution is, and the higher the number, the more basic the solution is. Neutral solutions (not acidic or basic) have a pH of 7.



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2 00010 M NaOH [OH-] =  
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